

PATENT SPECIFICATION

(11) 1 257 352

NO DRAWINGS

1 257 352

- (21) Application No. 20768/69 (22) Filed 23 April 1969
 (31) Convention Application No. 723 915 (32) Filed 24 April 1968 in
 (33) United States of America (US)
 (45) Complete Specification published 15 Dec. 1971
 (51) International Classification B 01 j 11/08; C 07 d 1/08
 (52) Index at acceptance
 B1E 20Y 25Y 303 304 308 321 322 32Y 331 33Y 341
 342 344 34Y 370 371 372 37Y 400 40Y 420 42Y
 510 550 55Y 690 692
 C2C 170—188—275 215 246 247 253 25Y 305 30Y 458 ZA
 (72) Inventor DENNIS ANTHONY DEMAIO



(54) PRODUCTION OF ETHYLENE OXIDE AND CATALYSTS FOR USE THEREIN

(71) We, HALCON INTERNATIONAL INC., of 2 Park Avenue, New York 16, New York, United States of America, a corporation organised and existing under the laws of the State of Delaware, United States of America, do hereby declare the invention, for which we pray that a patent may be granted to us, and the method by which it is to be performed, to be particularly described in and by the following statement:—

This invention relates to the production of ethylene oxide by the direct oxidation of ethylene and particularly to ethylene oxide catalysts and support material for such catalysts. The invention is also concerned with a process for the direct oxidation of ethylene to ethylene oxide wherein said oxidation is effected in the presence of such catalysts and in the absence of halogenated inhibitors.

In general, commercial production of ethylene oxide is accomplished by the vapour phase reaction of ethylene with molecular oxygen in the presence of a silver catalyst. The silver is usually in finely divided metallic form distributed on or throughout an inert support or carrier material. The latter are indicated in the prior art as including such materials as clay, bauxite, bentonite, kieselguhr, active carbon, pumice stone, silica and fused alumina. The manner of application of the silver to the support material and its effect on ethylene oxide productivity has been the subject of numerous investigations. In this regard, U.S. Patent No. 3,172,893 discloses a method of application wherein silver is coated onto the surfaces of support particles by spraying with a slurry of precipitated silver oxide. Patent No. 754,593 discloses another method of silver application wherein silver is impregnated throughout the pores of the carrier by immersion of the latter

in an aqueous silver solution, followed by drying and reduction to metallic silver.

Experience with the aforementioned silver catalysts has met with varying degrees of success, notwithstanding the close control of reaction conditions which the sensitive nature of the ethylene oxide process requires. Investigations have shown that results often vary depending on the method of silver deposition as well as the characteristics of the support material employed without a clear indication of the catalyst mechanism. The apparent complexity of such mechanism is compounded by the universal need for tempering the catalyst activity with a halogenated inhibitor; that is, a halogenated hydrocarbon capable of exerting an inhibiting effect upon the catalytic oxidation reaction such as ethylene dichloride. The latter generally comprises a portion of the gaseous feed mixture to the reaction zone and while small relative to the quantities of ethylene and oxygen consumed during production of ethylene oxide, nevertheless is a cost factor.

The economic penalty attendant upon the employment of a halogenated inhibitor in the ethylene oxide process, however, is not measured by this cost factor alone. The metering of such inhibitor requires very careful control and regulatory apparatus together with the expense necessarily incurred by expanded technical personnel requirements account a penalty significantly in excess of the cost of the inhibitor material alone. The presence of a halogenated inhibitor in the ethylene oxide process necessarily interferes with product purity, requiring additional steps to effect removal or separation. Process control suffers to some extent since an additional variable is introduced to the system. Of prime importance, however, is the fact that the presence of a halogenated

inhibitor necessitates the employment of more expensive materials of construction to withstand the corrosive nature of the material. Additional disadvantages may be attributed to the use of halogenated inhibitors including the necessity of taking neutralisation measures where high purity oxygen is employed to effect the partial oxidation of ethylene.

Accordingly, a principal object of the present invention is to provide a supported catalyst for use in the controlled, partial oxidation of ethylene to ethylene oxide which avoids the necessity of employing a halogenated inhibitor.

Another object of the invention is to provide a silver impregnated catalyst for use in the controlled oxidation of ethylene to ethylene oxide which obviates the need of a halogenated inhibitor and also permits of improved productivity and/or lower operating temperature.

The present invention is concerned essentially with a silver-containing catalyst for use in the production of ethylene oxide. The invention stems from the discovery that use of a support material which is porous in nature and characterized by a limited range of pore diameters, the average pore diameter of which falls in a narrow range, can eliminate the heretofore universal need for halogenated inhibitors to temper or otherwise control the activity of the silver-containing catalysts employed in the controlled, partial oxidation of ethylene to ethylene oxide. The invention contemplates that ethylene oxide catalysts prepared with the unique support material will be of the "impregnated" type that is, the catalyst is formed by filling the pores of the support material with a silver solution from which the silver solute is deposited. A discussion of an illustrative technique and procedure employed to prepare such a catalyst is provided in the aforementioned British patent 754,593.

The principles underlying the present invention are admittedly complex and the particular centres for catalyzing the different avenues of ethylene oxidation are not fully defined. Theoretically, some catalyst centers exhibit activity in the production of ethylene oxide from ethylene, while others exhibit activity in the formation of carbon dioxide from ethylene. It may also be conjectured that catalyst centers exist which effect the further oxidation of ethylene oxide to acid aldehydes and carbon dioxide. The picture is further complicated by the fact that realistic reaction rates and conversions attend high temperatures while low temperatures promote the selectivity of ethylene oxide. Normally, sufficiently high reaction temperatures are employed to provide an economically feasible ethylene conversion while a halogenated inhibitor, such as ethylene dichloride, is employed in the cycle gas to selectively poison those catalyst centers which exhibit undesirable activity, e.g., exhibit activity in the formation of carbon dioxide. While applicants do not intend to be bound

thereby, in principle, it would appear that catalyst centers exhibiting undesirable activity are minimized by the more homogeneous distribution of silver obtained by deposition thereof upon a porous support material wherein a substantial portion of the pores have diameters which fall within a limited range and wherein the average pore diameter falls within a narrow range.

For the purpose of the invention a support material may be regarded as porous when it has a minimum apparent porosity of 30%. It may be noted that apparent porosity may be determined by the mercury porosimeter method described by Drake and Ritter in "The Analytical Edition of Industrial Engineering Chemistry", Vol. 17, page 787 (1945).

A "substantial portion of the pores" may be regarded as embracing a minimum of 90% of the total pores of the support material.

The invention contemplates a "limited range of pore diameters" as including a maximum pore diameter range of 30 microns. Pore diameters may be determined by recourse to the B.E.T. nitrogen absorption method described by Brunauer, Emmett and Teller in "The Journal of the American Chemical Society", Vol. 60, page 309 (1938) and the Gurvicha equation described in "The Journal of the Russian Physical Chemical Society", Vol. 47, page 805, (1915). Thus the average pore diameter is defined by the ratio of the pore volume, as determined by the mercury porosimeter method, to the pore surface area, as determined by the B.E.T. nitrogen absorption method, in accordance with the Gurvicha equation.

The invention also contemplates that the "average pore diameter" will be of a size such that neither too low nor too high a diffusion rate is encountered in practice. All other factors being equal, the former is attended by poor ethylene oxide selectivity while the latter is attended by poor catalyst activity. It has been determined that an average pore diameter range of 4 to 10 microns in combination with the aforementioned catalyst support characteristics permits a diffusion rate giving the optimum combination of ethylene oxide selectivity and catalyst activity without the necessity of employing a halogenated inhibitor.

Thus, in accordance with the present invention, we provide a catalyst for use in the controlled partial oxidation of ethylene to ethylene oxide, said catalyst comprising a catalyst support impregnated with silver and said catalyst support comprising inert, porous particles having a minimum apparent porosity of 30% and wherein at least 90% of the pores have a maximum pore diameter range of 30 microns, the average pore diameter being in the range of 4 to 10 microns.

Practice of the ethylene oxide process using the catalyst of the invention and in the absence of halogenated inhibitor may be conducted with a considerable degree of flexibility. Operation

of the process may proceed at significantly reduced reaction temperatures while maintaining current productivity levels. Such operation may be dictated where maximum service life is a controlling factor. On the other hand, significantly increased productivity can be achieved without exceeding current operating temperatures.

In general, the invention catalyst support may comprise any porous, inert material which does not detrimentally influence the catalytic reaction wherein it is employed. Accordingly, in the catalytic oxidation of ethylene to ethylene oxide, suitable support materials include alpha alumina, silicon carbide, fused aluminum oxide, mixtures of alumina and silica. Desirably, the support material comprises 90% or more by weight alpha alumina and 1 to 6% by weight silica. A preferred support material comprises 90% or more by weight alpha alumina, 1 to 6% by weight silica and 0.1 to 0.4% by weight baryta.

The form of the inventive support may be described as a pellet having linear and diametrical dimensions suitably within the range of 1/8 to 5/16 inch and may take almost any solid geometrical configuration. From the standpoint of simplicity, such configuration is suitably cylindrical, spheroidal or spherical. It is recognized that engineering factors, such as ability to pack uniformly, mechanical strength, pressure drop and stability, may influence the choice of configuration and, accordingly, it is contemplated that more complex configurations, such as saddles and rings, may be dictated.

The catalyst supports may be prepared as follows: The support material, *e.g.*, high-purity aluminum oxide particles, preferably in the alpha alumina phase, are thoroughly mixed with temporary and permanent binders. The temporary binders are thermally decomposable organic compounds of moderate to high molecular weight (*i.e.*, molecular weights above about 250) and, on decomposition, produce the pore structure of the support. The permanent binders are inorganic clay-type materials having fusion temperatures below that of the alumina and are responsible for imparting mechanical strength to the finished support. Silica and baryta can also be added in quantity sufficient to obtain a finished support of the desired composition. After thorough dry-mixing, sufficient water is added to the mass to form the mass into a paste-like substance. The catalyst support particles are then formed from the paste by conventional means such as, for example, high pressure extrusion, granulation or other ceramic forming processes. The particles are then dried and are subsequently fired at an elevated temperature which is in the range of 120°C. to 1500°C.

In the firing step, the temporary binders are thermally decomposed to carbon dioxide and water and are volatilized, leaving voids in the

support mass. These voids are the genesis of the pore structure of the finished support. Suitable temporary binders include such materials as the celluloses and substituted celluloses, *e.g.*, cellulose itself, methyl cellulose, ethyl cellulose, and carboxyethyl cellulose, stearates such as organic stearate esters, *e.g.*, methyl or ethyl stearate, or waxes. As firing is continued, the temperature reaches the point at which the permanent binder (inorganic clay such as the kaolins or the ball clays) fuses. The catalyst support is then permitted to cool and, during cooling, the permanent binder sets and acts as a cement to bond the catalyst support particles and thereby impart mechanical strength to the support and ensure maintenance of the pore structure.

Catalyst supports of desired characteristics can be readily produced by the foregoing procedure. Control of pore size, pore size distribution and porosity are readily effected by appropriate adjustment in known manner of the size of the starting alumina particles, and of the particle size and concentration of the temporary and of the permanent binders in the starting mix. The larger the starting alumina particle size, the greater will be the porosity of the finished catalyst. The more homogeneous in size are the alumina particles, the more uniform will be the pore structure. Similarly, increasing the concentration of the concentration of the temporary binder will also increase the overall porosity of the finished catalyst support.

As aforementioned, the support material of the present invention is intended for use in the preparation of ethylene oxide catalysts of the impregnated type rather than the coated type. Impregnation may suitably be accomplished by immersion of the support material in a silver solution, *e.g.*, an aqueous solution of silver nitrate or silver lactate. Essentially, the solution with the support material immersed therein is heated, desirably to boiling, to effect degasification of the support material and to insure complete and uniform impregnation thereof. The impregnated support material may then be removed for the excess solution, and carefully and slowly dried.

Deposited silver salt is then converted to a catalytically active form, *e.g.*, by hydrogen reduction in a case such as silver nitrate or by thermal decomposition in a case such as silver lactate.

The amount of silver deposited upon the support material will be controlled by the concentration of the silver salt in the impregnating solution and by the porosity of the catalyst support. The actual amount of silver on the support after reduction or thermal decomposition can vary from about 5% to about 30% by weight of the reduced or thermally decomposed impregnated catalyst, the minimum value being dictated by insufficient activity while the maximum value represents

the point where the catalyst begins to resemble a coated catalyst in physical properties.

Promoters are often employed in ethylene oxide catalysts in combination with silver catalytic material for the primary purpose of extending catalyst service life while also influencing particular aspects of the ethylene oxide reaction. In this regard, materials comprising barium, copper, mercury or tin are illustrative of promoters commonly employed in the catalytic oxidation of ethylene to produce ethylene oxide and contemplated by the present invention. Such materials are generally introduced to the silver catalytic material-containing media employed to impregnate the support and thus are generally co-deposited with the silver catalytic material.

In practice, a water-soluble salt of the promoting metal is added to the silver salt solution in such proportions that the concentration of promoter metal (based on the equivalent metal content of the oxide) is from about 0.1% to about 3.0% by weight of the silver metal content of the reduced catalyst. Amounts of

promoter metal less than about 0.1% of the silver generally exert a negligible effect on catalyst life while amounts higher than about 3% lead to reactions which are difficult to control. The preferred concentration of promoter metal is from 0.1% to 1.5% by weight of the silver metal. The water-soluble salt of the promoter metal, of course, should be consonant with the technique employed to treat the deposited silver salt. Thus, in the case of reduction, the salt should be reducible to the oxide at the catalyst reduction temperatures employed. For this purpose, the nitrite and nitrate salts are suitable. In the case of thermal decomposition, the salt should be thermally decomposable at the thermal decomposition temperatures employed. For this purpose, the lactate and oxalate salts are suitable.

In the course of the investigation leading to the present invention, 3/16 inch spherical catalyst support pellets were prepared using the procedure set forth on page 6 and having the following chemical and physical properties:

TABLE I

Chemical Composition	A	B	C	D
Al ₂ O ₃	86.96	88.12	86.20	93.50
SiO ₂	11.65	9.79	12.40	5.30
Fe ₂ O ₃	0.30	0.20	0.24	0.18
TiO ₂	0.42	0.16	0.46	0.11
CaO	0.05	0.55	0.06	0.30
MgO	0.11	0.43	0.03	0.25
Na ₂ O	0.32	0.19	0.34	0.07
K ₂ O	0.17	0.07	0.20	0.04
ZrO ₃ + HfO ₂	0.02		0.04	
BaO		0.49		0.27
MnO			0.01	
Physical Properties				
Apparent Porosity, %	51	56	40	53
Pore Diameter Range, μ	4—70	1—40	1—2	1—3.5
Average Pore Diameter, μ	19	7	1.8	5

The pellets were then impregnated with silver by immersion in a silver-containing solution prepared as follows: 1000 parts by weight of silver oxide is added slowly with stirring to 1224 parts by weight of an 85% solution of lactic acid in water over a period of 30 to 45 minutes, the mixture being continuously cooled to avoid temperatures in excess of 95°C. and thereafter 100 parts by weight of a 15% solution of hydrogen peroxide in water is slowly added to produce a clear yellow solution, whereupon 39 parts by weight of a 44.4% barium lactate solution in water is added. The silver-containing solution is maintained at a temperature of 90 to 95°C. during immersion of the pellets, which are permitted to soak for 5 to 15 minutes with occasional stirring. The pellets are then separated from the solution by draining. After about 15 minutes of draining, the pellets are dried for at least 10 hours at a temperature of 60°C to 70°C in an air atmosphere. The dried pellets are then heated gradually over a period of about 4 hours to a temperature of about 250°C. in an air atmosphere and maintained at such temperature for an additional 4 hours to effect activation.

Evaluation of the catalyst pellets, prepared

as aforescribed, was conducted at primary reactor conditions in a pressure reactor system for the controlled oxidation of ethylene to ethylene oxide. Feed gas comprising 5.0% ethylene, 6.0% oxygen, 0.1% ethane, balance essentially nitrogen, was passed at a flow rate of 53.5 lbs/hr/tube and a pressure of 315 psia through the prepared catalyst at a bed height of 23 ft 5 in. The results in terms of ethylene oxide selectivity and productivity obtained at recited reactor temperatures and halogenated inhibitors contents are set forth in Table II. It may be noted at this point that the term "selectivity" is commonly employed to indicate the percentage of converted ethylene constituting ethylene oxide, such value being of controlling importance in the operation of a primary or main reactor. The term "productivity" has a more general connotation but may be regarded simply as the difference between product output and product input across the reactor. In this regard, ethylene oxide productivity is commonly measured indirectly by Δ^{EO} , i.e., the difference between the ethylene oxide concentrations in the reactor effluent and inlet gases.

TABLE II

	A	B	C	D	
Reactor temperature, °C.	245	235	245	220	230
Ethylene Dichloride, ppm	0.03	0.04	0	0	0
Selectivity	70.5	68.3	70.8	74.9	70.4
Δ	1.16	0.03	0.96	1.32	1.3

Analysis of the data of Table II indicates quite clearly that catalysts prepared with support D, i.e., the inventive support, permit the operation of the ethylene oxide process at significantly lower reaction temperatures, e.g. 220° to 230°C, without the necessity of employing a halogenated inhibitor. Comparison of inventive support D with support C, the only comparative support not requiring utilization of a halogenated inhibitor, reveals a most striking improvement in reaction temperature and productivity. Table II reveals that inventive support D may be employed at a reactor temperature 25°C lower than in the case of inventive support C with little reduction in productivity, and at a temperature of 15°C. lower with a substantial increase in productivity.

Having determined the new results and advantages of the invention under primary reactor

conditions, comparative data was sought at purge reactor conditions. Accordingly, feed gas comprising 1.4% ethylene, 6.0% oxygen, 7.0% carbon dioxide, 0.001% ethane, balance essentially nitrogen, was passed at a flow rate of 37.5 lbs/hr/tube and a pressure of 315 psia through beds of catalysts prepared with support pellets A, B and D. No halogenated inhibitor was used. The results in terms of ethylene conversion and ethylene oxide productivity are given for different reactor temperatures in Table III. The term "conversion" is employed to indicate the percentage of ethylene which is converted, i.e., without regard to the products resulting from such conversion. As a purge reaction generally represents the last opportunity to wrest product from the raw materials, a high conversion value is of utmost importance.

TABLE III

Temperature, °C.	A		B		D	
	Δ^{EO}	Conv. %	Δ^{EO}	Conv. %	Δ^{EO}	Conv. %
235				57.2	0.80	80.3
240					0.84	86.0
245	0.34	29.6	0.56	66.1		
250					0.82	88.0
255	0.48	45.0				
265	0.56	55.7	0.64	73.5		

5 Recourse to Table III reveals most dramatically that the catalyst prepared with support pellets D and illustrative of the invention permits the obtention of highest conversion and productivity values and at lowest reactor temperatures. It is apparent that significant economic advantages will attend practice of the invention since the data of Table III attests to a minimum ethylene recycle to the purge reactor as well as a maximum life of the catalyst employed therein.

WHAT WE CLAIM IS:—

15 1. A catalyst for use in the controlled partial oxidation of ethylene to ethylene oxide, said catalyst comprising a catalyst support impregnated with silver and said catalyst support comprising inert, porous particles having a minimum apparent porosity of 30% and
20 wherein at least 90% of the pores have a maximum pore diameter range of 30 microns, the average pore diameter being in the range of 4 to 10 microns.

25 2. A catalyst according to claim 1 wherein said particles comprise silica-alumina.

3. A catalyst according to claim 2 wherein said silica is in the range of 1 to 6% by weight.

4. A catalyst according to claim 2 wherein said particles comprise 0.1 to 4.0% by weight baryta.

5. In a process for the controlled, partial oxidation of ethylene to ethylene oxide, the improvement comprising carrying out said oxidation in the presence of a catalyst according to any of claims 1 to 4 and in the effective absence of a halogenated inhibitor.

6. A catalyst according to claim 1 substantially as hereinbefore described and exemplified.

7. A process according to claim 5 substantially as hereinbefore described and exemplified.

8. Ethylene oxide when produced by a process according to claim 5 or 7.

MATHYS & SQUIRE,
Chartered Patent Agents,
10 Fleet Street,
London, E.C.4.
Agents for the Applicants.